

plane *B* to plane *D* = 130.4°) and a pseudo-symmetry plane through C(16). The oxygen atoms in the molecule are bound by intramolecular hydrogen bonds. Inter-molecular van der Waals forces hold the molecules together in the structure.

One of us (P.N.) is a recipient of a Deutscher Akademischer Austauschdienst scholarship, for which he is grateful.

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The Crystal and Molecular Structure of Hydrotris(1-pyrazolyl)boratobenzenediazodicarbonyl Molybdenum HB(pz)₃Mo(CO)₂NNC₆H₅

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Hydrotris(1-pyrazolyl)boratobenzenediazodicarbonyl molybdenum, HB(C₃N₂H₃)₃Mo(CO)₂N₂C₆H₅, crystallizes in the space group *P*1̄. The unit-cell constants are: *a* = 9.495 ± 0.010, *b* = 11.878 ± 0.014, *c* = 9.190 ± 0.009 Å; α = 102° 56' ± 5', β = 85° 56' ± 3', γ = 100° 44' ± 4', *Z* = 2; *D_x* = 1.59, *D_{meas}* = 1.59 ± 0.01. Data for 2571 reflections were collected at room temperature on a Picker automated diffractometer. The crystal structure was solved by the Patterson method and refined by block-matrix least-squares calculations on 2375 observed reflections to a reliability index of 0.038. Molybdenum and boron lie on the pseudo threefold axis of the molecule and are bonded to adjacent nitrogen atoms on each of the three pyrazole rings. The average B–N bond distance is 1.55 Å. The average Mo–N(pyrazolyl) bond distance is 2.22 Å whereas the Mo–N(diazo) bond distance is 1.83 Å. The geometry about the Mo atom is that of a slightly distorted octahedron. The boron atom exhibits slightly distorted tetrahedral geometry.

Introduction

In the last few years a number of transition metal pyrazolyl borate complexes have been prepared (Trofimenko, 1967, 1968, 1969). In order to gain more insight into the stereochemistry of such compounds we have undertaken the structural determination of the pyrazolyl borate complex hydrotris(1-pyrazolyl)boratobenzenediazodicarbonyl molybdenum.

A similar compound: hydrotris(1-pyrazolyl)boratodiazodicarbonylnitrosyl molybdenum, HB(pz)₃Mo(CO)₂NO, has been recently studied (Holt, 1970). However, in the nitrosyl complex a statistical disorder involving

the carbonyl and nitrosyl groups has been found. In addition the presence of a disordered solvent molecule has prevented the successful refinement of the structure. The present complex does not show a disordering of the carbonyl and benzenediazo groups because of the very different steric encumbrance of these ligands.

Experimental

Red-orange crystals of the complex were provided by S. Trofimenko. The one chosen was a plate of dimensions 0.45 × 0.24 × 0.04 mm. Weissenberg photographs showed that the crystal was triclinic. Centrosymmetry was assumed and confirmed by the subsequent refinement.

The crystal was mounted on a Picker automated diffractometer with the *a* axis parallel to the spindle axis. The unit-cell constants were determined from a least-squares fit of the angular positions of twelve independent reflections. Data were taken to a 2θ value

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of 45° with Zr-filtered Mo $K\alpha$ radiation. No attenuators were used. A take-off angle of 3° , a 2θ scan mode and a scan rate of 2° per minute were used for the data collection. The large scan range was chosen because some of the peaks were nearly 1.5° wide. The peak was scanned for one hundred seconds. Backgrounds on either side of the peak were measured for ten seconds. Data were collected for 2571 reflections of which 2375 were observed. Two standard reflections were measured every forty reflections. The standard deviation of the intensities of these two reflections was $\pm 1.5\%$.

Crystal data

Space group $P\bar{1}$

$a = 9.496 \pm 0.011 \text{ \AA}$ M.W. 472.12

$b = 11.877 \pm 0.014$ Cell volume 987 \AA^3

$c = 9.189 \pm 0.009$ Density, observed 1.59 ± 0.01
X-ray 1.59 g.cm^{-3}

$\alpha = 102^\circ 56' \pm 5'$

$\beta = 85^\circ 56' \pm 1'$

$\gamma = 100^\circ 44' \pm 4'$

$\mu(\text{Mo } K\alpha) 6.9 \text{ cm}^{-1}$

The background was subtracted and each peak was corrected for Lorentz and polarization factors. A reflection was considered observed if the total integrated intensity minus the background was greater than 200 counts for a 100 second scan. No absorption corrections were applied. The minimum value of μR is 0.042 at $2\theta = 0$ perpendicular to the plane of the crystal plate. The maximum value of μR is 0.110 at $2\theta = 40^\circ$ parallel to the plane of the crystal plate.

Determination of the structure and refinement

The structure factors were phased by the heavy atom method. The first Fourier synthesis located all 29 non-hydrogen atoms in the asymmetric unit. The atomic positions and isotropic thermal parameters were refined by six cycles of block-matrix least-squares calculations (Immirzi, 1967) to an R value of 0.079. Three cycles including anisotropic thermal parameters reduced the R value to 0.049.

At this point a difference synthesis showed the positions of the fifteen hydrogen atoms of the asymmetric

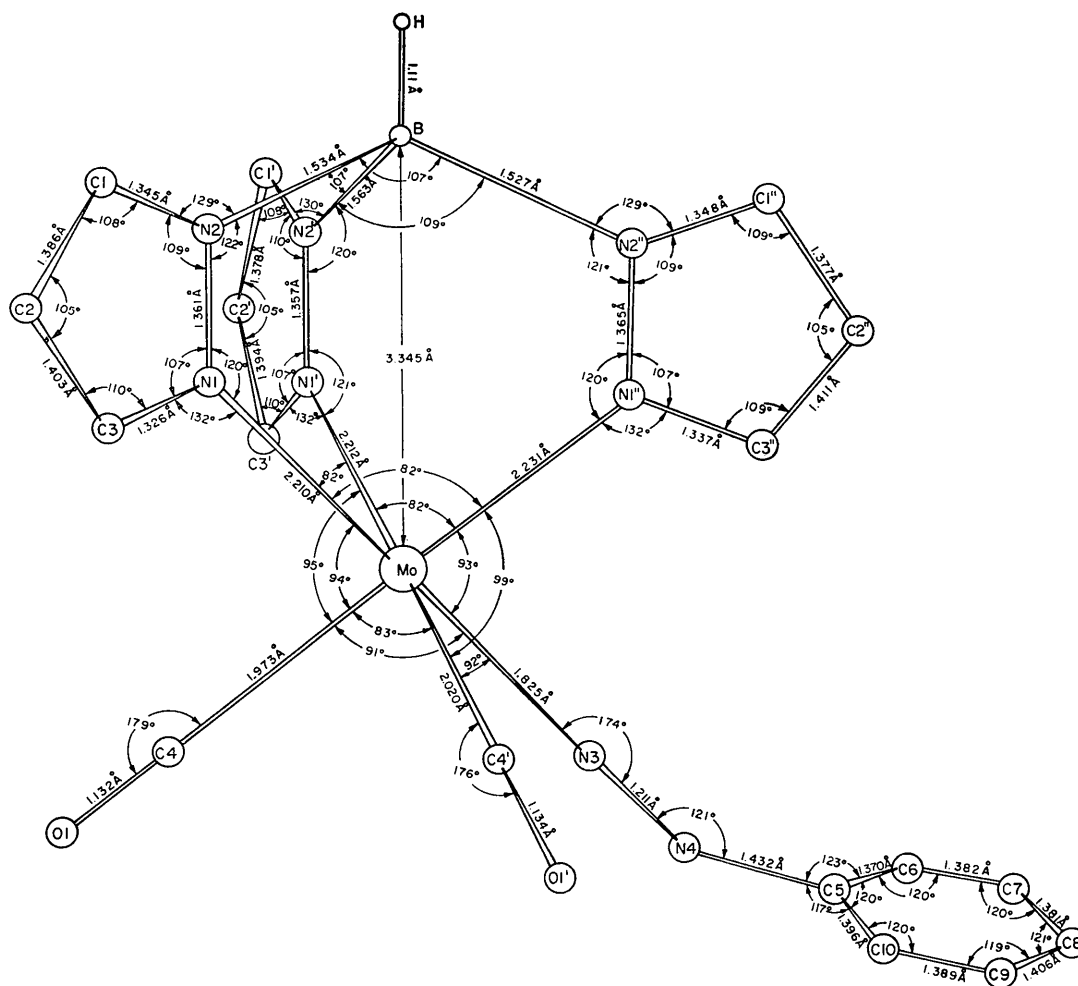


Fig. 1. Molecular geometry of hydrotris(1-pyrazolyl)boratobenzenediazocarbonyl molybdenum. The most important conformational parameters are shown.

unit. The hydrogen atoms were included in the next refinement with isotropic thermal factors. The R index dropped to 0.038. The function minimized was: $\sum(w|F_e - KF_o|)^2$. The form of Cruickshank's (1949) weighting function was $w = (0.77777 + 0.01243 F_{o\min} + 1.9888 \times 10^{-4} F_o^2)_{\max}$. The standard deviation of each reflection was used as a basis for the weighting function in the earlier stages of the refinement. The list of observed and calculated structure factors is given in Table 1. The atomic scattering factors were taken from Moore (1963). No corrections were included for anomalous scattering.

The final positional and thermal parameters are listed in Table 2 for non-hydrogen atoms. The anisotropic thermal parameters for the non-hydrogen atoms are listed in Table 3. The positional and isotropic thermal parameters for the hydrogen atoms are listed in Table 4. The standard deviations for the positional parameters of the hydrogen atoms are approximately ten times greater than those for the carbon atoms. The standard deviations of the thermal parameters for the hydrogen atoms are approximately $\pm 1 \text{ \AA}^2$.

Table 2. Positional parameters in fractional coordinates

	x	y	z
O(1')	-0.0441 (5)	-0.0462 (4)	0.3031 (5)
N(1)	0.2860 (5)	0.0562 (4)	0.0608 (5)
N(2)	0.3370 (5)	0.1061 (4)	-0.0560 (5)
N(1')	0.4014 (4)	0.2938 (4)	0.2298 (4)
N(2')	0.4393 (5)	0.3089 (4)	0.0896 (5)
N(1'')	0.1146 (5)	0.2388 (4)	0.1084 (5)
N(2'')	0.1933 (5)	0.2651 (4)	-0.0132 (5)
N(3)	0.1417 (4)	0.2620 (4)	0.4333 (5)
N(4)	0.1063 (5)	0.3351 (4)	0.5382 (5)
C(1)	0.3771 (6)	0.0235 (6)	-0.1702 (7)
C(2)	0.3514 (7)	-0.0829 (6)	-0.1267 (7)
C(3)	0.2955 (6)	-0.0570 (5)	0.0204 (7)
C(1')	0.5618 (7)	0.3904 (6)	0.0941 (8)
C(2')	0.6052 (7)	0.4293 (6)	0.2408 (8)
C(3')	0.5014 (6)	0.3675 (5)	0.3226 (7)
C(1'')	0.1160 (7)	0.3137 (5)	-0.0902 (7)
C(2'')	-0.0147 (7)	0.3210 (6)	-0.0172 (8)
C(3'')	-0.0111 (6)	0.2729 (5)	0.1084 (7)
C(4)	0.3057 (6)	0.0957 (5)	0.4026 (6)
C(4')	0.0447 (6)	0.0293 (5)	0.2874 (6)
C(5)	-0.0423 (6)	0.3372 (5)	0.5755 (6)
C(6)	-0.0729 (7)	0.4413 (5)	0.6608 (7)
C(7)	-0.2133 (7)	0.4500 (6)	0.7045 (7)
C(8)	-0.3229 (7)	0.3533 (7)	0.6631 (7)
C(9)	-0.2930 (7)	0.2459 (7)	0.5755 (9)
C(10)	-0.1514 (7)	0.2382 (6)	0.5326 (8)
B	0.3440 (7)	0.2378 (6)	-0.0470 (7)

Standard deviations are given in parentheses.

	x	y	z
Mo	0.2091 (1)	0.1638 (1)	0.2708 (1)
O(1)	0.3593 (5)	0.0563 (5)	0.4792 (5)

Description of the structure

The most important conformational parameters of the molecule are listed in Table 5 and Fig. 1.

Table 3. Thermal parameters for the non-hydrogen atoms in the form

$$\exp[-\frac{1}{4}(B_{11}a^2h^2 + B_{22}b^2k^2 + B_{33}c^2l^2 + 2B_{12}a^*b^*hk + 2B_{13}a^*c^*hl + 2B_{23}b^*c^*kl)]$$

	B_{11}	B_{22}	B_{33}	B_{12}	B_{13}	B_{23}
Mo	2.866 (0.019)	3.051 (0.020)	2.546 (0.019)	0.067 (0.015)	0.017 (0.015)	0.449 (0.015)
O(1)	4.899 (0.227)	6.784 (0.268)	5.669 (0.251)	1.879 (0.202)	-0.389 (0.187)	2.844 (0.212)
O(1')	5.045 (0.236)	3.953 (0.208)	6.122 (0.259)	-0.514 (0.178)	0.624 (0.196)	0.413 (0.185)
N(1)	3.279 (0.202)	2.806 (0.187)	2.555 (0.187)	0.835 (0.157)	0.067 (0.151)	-0.055 (0.150)
N(2)	2.800 (0.190)	4.159 (0.218)	2.305 (0.185)	0.999 (0.167)	-0.039 (0.147)	0.176 (0.159)
N(1')	2.987 (0.190)	2.479 (0.178)	2.413 (0.178)	0.396 (0.150)	-0.077 (0.144)	0.425 (0.142)
N(2')	2.684 (0.188)	3.533 (0.207)	3.398 (0.209)	0.433 (0.158)	0.146 (0.156)	1.116 (0.167)
N(1'')	2.853 (0.191)	3.235 (0.199)	3.048 (0.199)	0.763 (0.158)	-0.361 (0.153)	0.682 (0.158)
N(2'')	3.587 (0.208)	3.349 (0.205)	2.539 (0.189)	0.754 (0.167)	-0.412 (0.156)	0.695 (0.158)
N(3)	2.541 (0.177)	2.886 (0.184)	2.266 (0.175)	0.581 (0.146)	-0.129 (0.138)	-0.238 (0.142)
N(4)	3.226 (0.203)	3.070 (0.199)	3.271 (0.204)	0.441 (0.163)	0.131 (0.160)	0.095 (0.161)
C(1)	3.279 (0.258)	5.145 (0.313)	3.227 (0.260)	1.517 (0.232)	-0.123 (0.203)	-0.777 (0.228)
C(2)	4.014 (0.285)	4.208 (0.286)	3.968 (0.287)	1.258 (0.231)	-1.005 (0.227)	-1.123 (0.228)
C(3)	3.029 (0.243)	3.014 (0.244)	4.322 (0.284)	0.286 (0.197)	-0.719 (0.207)	-0.407 (0.209)
C(1')	3.511 (0.272)	4.319 (0.294)	5.630 (0.343)	0.443 (0.228)	0.399 (0.240)	2.090 (0.260)
C(2')	3.791 (0.188)	3.759 (0.207)	5.627 (0.209)	-0.241 (0.158)	-0.713 (0.156)	0.764 (0.167)
C(3')	3.817 (0.263)	3.325 (0.247)	3.686 (0.266)	0.658 (0.207)	-0.950 (0.209)	0.608 (0.206)
C(1'')	4.668 (0.301)	3.801 (0.265)	3.598 (0.272)	0.554 (0.228)	-1.642 (0.229)	1.286 (0.218)
C(2'')	4.276 (0.297)	4.063 (0.284)	4.990 (0.319)	1.112 (0.238)	-1.421 (0.243)	0.658 (0.242)
C(3'')	2.504 (0.230)	3.683 (0.253)	4.467 (0.288)	0.641 (0.197)	-0.701 (0.204)	0.397 (0.216)
C(4)	3.226 (0.243)	3.984 (0.263)	3.094 (0.243)	0.773 (0.204)	-0.305 (0.190)	0.956 (0.203)
C(4')	3.855 (0.265)	3.543 (0.257)	3.239 (0.251)	0.858 (0.213)	0.274 (0.203)	0.252 (0.203)
C(5)	2.967 (0.228)	3.440 (0.237)	2.489 (0.218)	0.818 (0.190)	0.076 (0.177)	0.635 (0.184)
C(6)	4.415 (0.296)	3.231 (0.256)	3.832 (0.281)	0.703 (0.221)	0.460 (0.222)	0.022 (0.213)
C(7)	5.307 (0.335)	4.353 (0.302)	4.323 (0.302)	2.102 (0.263)	0.499 (0.250)	0.780 (0.241)
C(8)	4.125 (0.296)	7.028 (0.392)	3.487 (0.285)	2.576 (0.284)	0.835 (0.230)	1.534 (0.266)
C(9)	3.859 (0.303)	5.315 (0.344)	6.273 (0.389)	0.780 (0.260)	0.200 (0.269)	0.904 (0.293)
C(10)	3.924 (0.290)	4.157 (0.295)	5.033 (0.328)	0.908 (0.236)	0.673 (0.240)	-0.281 (0.247)
B	3.247 (0.285)	4.366 (0.314)	2.826 (0.268)	0.472 (0.241)	0.061 (0.216)	0.809 (0.234)

Table 4. *Positional and thermal parameters for the hydrogen atoms*

	<i>x</i>	<i>y</i>	<i>z</i>	<i>B</i>
H(C1)	0.4125 (78)	0.0425 (66)	-0.2649 (86)	4.3 (0.9) Å ²
H(C2)	0.3673 (69)	-0.1614 (56)	-0.1857 (72)	2.6 (0.7)
H(C3)	0.2629 (76)	-0.1079 (61)	0.0861 (81)	3.5 (0.8)
H(C1')	0.6064 (64)	0.4158 (52)	0.0070 (68)	1.7 (0.6)
H(C2')	0.6858 (21)	0.4903 (61)	0.2832 (87)	4.6 (0.9)
H(C3')	0.3857 (75)	0.2599 (61)	-0.1543 (79)	3.6 (0.8)
H(C1'')	0.1519 (73)	0.3339 (60)	-0.1824 (78)	3.2 (0.8)
H(C2'')	-0.0910 (81)	0.3576 (67)	-0.0334 (86)	4.3 (0.9)
H(C3'')	-0.0886 (67)	0.2591 (55)	0.1902 (71)	2.6 (0.7)
H(C6)	0.0088 (80)	0.5184 (65)	0.6953 (85)	4.6 (0.9)
H(C7)	-0.2384 (82)	0.5336 (66)	0.7654 (87)	4.9 (0.9)
H(C8)	-0.4201 (78)	0.3575 (64)	0.6964 (84)	4.2 (0.9)
H(C9)	-0.3676 (90)	0.1732 (73)	0.5455 (96)	5.6 (1.0)
H(C10)	-0.1293 (75)	0.1614 (60)	0.4674 (80)	3.6 (0.8)
H(B)	0.4923 (81)	0.3681 (67)	0.4237 (83)	3.6 (0.9)

Table 5. *Conformational parameters of the structural unit*

Standard deviations are given in parentheses.
Bond distances

Mo—N(1)	2.210 (4) Å
Mo—N(1')	2.212 (4)
Mo—N(1'')	2.231 (5)
Mo—N(3)	1.825 (4)
Mo—C(4)	1.973 (6)
Mo—C(4')	2.020 (6)
B—N(2)	1.534 (9)
B—N(2')	1.563 (8)
B—N(2'')	1.525 (8)
N(1)—N(2)	1.361 (6)
N(2)—C(1)	1.345 (8)
C(1)—C(2)	1.386 (10)
C(2)—C(3)	1.403 (9)
C(3)—N(1)	1.326 (7)
N(1')—N(2')	1.357 (6)
N(1')—C(3')	1.346 (7)
C(2')—C(3')	1.394 (9)
C(1')—C(2')	1.378 (10)
N(1'')—N(2'')	1.365 (6)
C(2'')—C(3'')	1.411 (9)
N(2'')—C(1'')	1.348 (8)
N(3)—N(4)	1.211 (6)
N(4)—C(5)	1.432 (7)

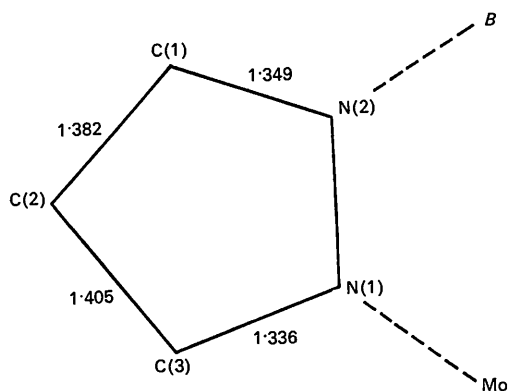


Fig. 2. Representation of a pyrazolyl ring in the molecule. Average parameters of the three pyrazolyl rings are given which shows that the twofold symmetry is lacking in the rings.

Table 5 (cont.)

C(5)—C(10)	1.396 (9)
C(9)—C(10)	1.389 (10)
C(8)—C(9)	1.406 (11)
C(7)—C(8)	1.381 (10)
C(6)—C(7)	1.382 (9)
O(1')—C(4')	1.134 (8)
O(1)—C(4)	1.132 (8)
C(1'')—C(2'')	1.377 (9)
N(2')—C(1')	1.350 (8)
N(1'')—C(3'')	1.337 (7)
C(5)—C(6)	1.370 (8)

Mo—B (non-bonded distance) = 3.345 (7)

Bond angles

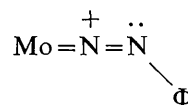
Mo—C(4')—O(1')	176.34 (15)°
Mo—C(4)—O(1)	179.05 (62)
Mo—N(3)—N(4)	174.21 (12)
Mo—N(1)—C(3)	131.81 (07)
Mo—N(1)—N(2)	120.51 (06)
Mo—N(1')—N(2')	121.24 (06)
Mo—N(1')—C(3')	131.90 (06)
Mo—N(1'')—N(2'')	120.43 (06)
N(1)—Mo—N(1')	82.20 (08)
N(1)—Mo—N(1'')	81.60 (08)
N(1')—Mo—N(1'')	81.66 (08)
N(1)—Mo—C(4)	94.49 (12)
N(1')—Mo—C(4)	95.36 (09)
N(1')—Mo—C(4')	172.44 (79)
N(1'')—Mo—C(4')	99.08 (09)
C(4)—Mo—C(4')	83.41 (13)
C(4')—Mo—N(3)	91.78 (12)
C(4)—Mo—N(3)	90.80 (09)
N(1')—Mo—N(3)	95.70 (10)
N(1'')—Mo—N(3)	93.03 (11)
B—N(2)—N(1)	121.78 (26)
B—N(2)—C(1)	128.76 (29)
B—N(2')—N(1')	120.58 (23)
B—N(2')—C(1')	129.67 (28)
B—N(2'')—N(1'')	121.29 (24)
B—N(2'')—C(1'')	129.29 (28)
N(2)—B—N(2')	107.07 (21)
N(2)—B—N(2'')	108.88 (22)
N(2')—B—N(2'')	107.38 (21)
N(1)—N(2)—C(1)	109.46 (24)
N(2)—C(1)—C(2)	108.20 (26)
C(1)—C(2)—C(3)	104.99 (26)
C(2)—C(3)—N(1)	109.66 (25)
C(3)—N(1)—N(2)	107.68 (21)
N(2')—N(1')—C(3')	106.84 (23)
N(1')—C(3')—C(2')	109.83 (25)
C(3')—C(2')—C(1')	105.24 (28)

Table 5 (cont.)

C(2')—C(1')—N(2')	108.34 (27)
C(1')—N(2')—N(1')	109.74 (21)
N(2'')—N(1'')—C(3'')	107.47 (22)
N(1'')—C(3'')—C(2'')	109.35 (24)
C(3'')—C(2'')—C(1'')	105.08 (27)
C(2'')—C(1'')—N(2'')	108.68 (25)
C(1'')—N(2'')—N(1'')	109.41 (21)
N(3)—N(4)—C(5)	121.09 (21)
N(4)—C(5)—C(10)	116.44 (27)
C(10)—C(9)—C(8)	118.99 (31)
C(9)—C(8)—C(7)	120.49 (32)
C(8)—C(7)—C(6)	119.97 (30)
C(7)—C(6)—C(5)	120.10 (27)
Torsion angle	
N(3)—N(4)—C(5)—C(6)	159.75

The coordinations of the molybdenum and boron atoms are slightly distorted octahedral and tetrahedral respectively. The hydrotris(1-pyrazolyl)borate ligand has nearly C_{3v} symmetry as foreseen by Trofimenko (1969). The average B—N bond distance is 2.22 Å, whereas the Mo—N(diazo) bond distance is 1.83 Å. The bond distances Mo—N(pyrazolyl) have nearly the value expected for neutral nitrogen atoms bonded to Mo. In fact the formal charge of these nitrogen atoms is only $-\frac{1}{3}$ because the uninegative charge on the $\text{HB}(\text{pz})_3^-$ ion must be shared among three nitrogen atoms. The shorter Mo—N(diazo) bond distance may

be explained in terms of a substantial contribution of the resonance structure



which also justifies the collinearity of the Mo—N—N atoms. The torsion angle around the bond N(4)—C(5) is approximately 20° ; a similar torsion angle has been found for *trans*-azobenzene ($\sim 17^\circ$) (DeLeange, Robertson & Woodward, 1939).

The conformation of each pyrazole ring is not characterized by C_2 symmetry and is in agreement with resonance between the two limiting structures *A* and *B*, with *A* given double weight:

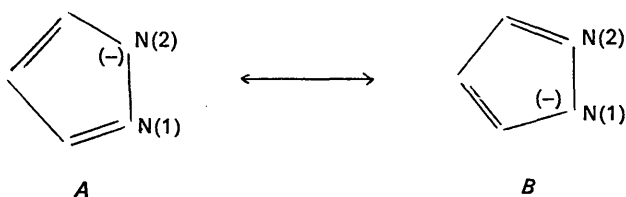


Fig. 2 shows the geometrical parameters of the pyrazolyl ligand averaged among all the three rings. Fig. 3 shows a projection of the structure on (001). The shortest intermolecular contact distances are reported.

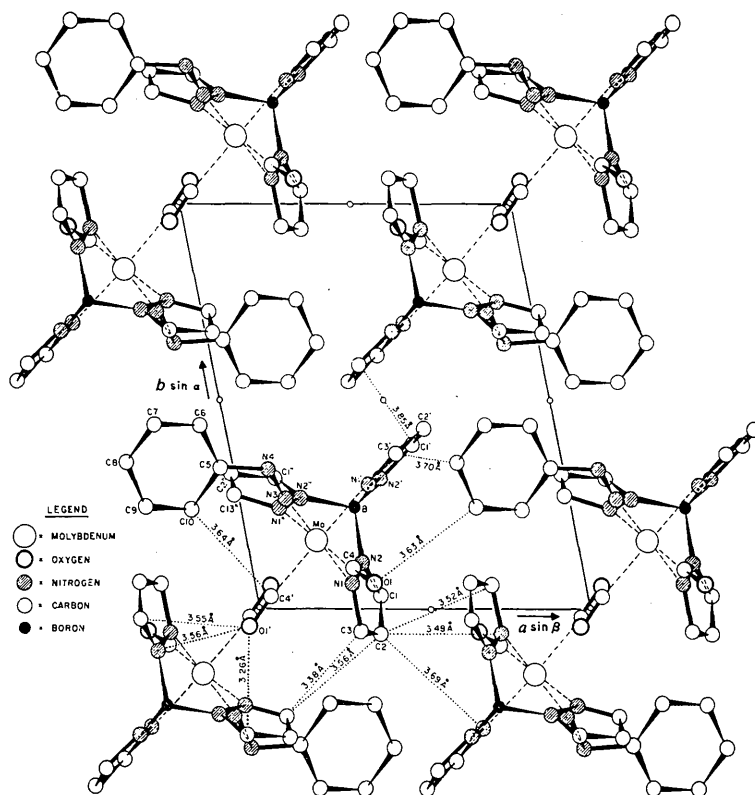


Fig. 3. Projection of the packing on (001). The shortest intermolecular contact distances are reported.

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Double Oxides Containing Niobium, Tantalum, or Protactinium.

IV. Further Systems Involving Alkali Metals

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The binary systems $A_2O-B_2O_5$ (where $A = \text{Rb}$ or Cs and $B = \text{Nb}$ or Ta) and $A_2O-Pa_2O_5$ (where $A = \text{Li}$, Na or K) have been partially examined. Unit-cell data are reported for most of the compounds prepared. In the $\text{Na}_2\text{O}-\text{Pa}_2\text{O}_5$ system, NaPaO_3 having a rhombohedrally distorted perovskite structure, and NaPa_3O_8 with a hexagonal unit cell have been observed. A value of 0.86 \AA has been calculated for the ionic radius of protactinium from the cell parameters of KPaO_3 and RbPaO_3 . The formation of APaO_3 compounds is discussed.

Introduction

Double oxides containing niobium or tantalum and alkali metals lithium to potassium, and some of those with caesium were reported on in the first paper of the present series (Whiston & Smith, 1965). This study is extended in the present paper to include the remaining rubidium and caesium compounds with niobium and tantalum, and also those of protactinium with the lighter alkali metals. These latter compounds have also been extensively studied by Keller in recent years.

Experimental

Preparative and X-ray techniques were outlined in part I (Whiston & Smith, 1965) and part III (Iyer & Smith, 1967). Fuller details of the experimental conditions are given by Iyer (1965).

The system $\text{Rb}_2\text{O}-\text{Nb}_2\text{O}_5$

Reisman & Holtzberg (1960) identified eight intermediate phases in the system $\text{Rb}_2\text{O}-\text{Nb}_2\text{O}_5$. The compositions and melting points are given in Table 1. None of the complex X-ray powder diffraction patterns were indexed by these authors. We have not been able to prepare single crystals of rubidium-rich phases (including RbNbO_3) for X-ray examination because of their

highly hygroscopic nature. We obtained crystals of ' $\text{Rb}_4\text{Nb}_6\text{O}_{17}$ ' and the lattice was found to be orthorhombic with

$$a = 6.42, b = 7.68, c = 38.55 \text{ \AA}.$$

The density measured (4.45 g.cm.^{-3}) does not agree with the calculated value (4.09 g.cm.^{-3}) with $Z = 4$ for the above composition. However the calculated density (4.56 g.cm.^{-3}) for $\text{Rb}_8\text{Nb}_{14}\text{O}_{39}$ with $Z = 2$ agrees quite well with the measured density.

Table 1. *Compositions and melting points*

Compound	Melting point
$\text{Rb}_4\text{Nb}_3\text{O}_{77}$	1351 °C (incongruent)
$\text{Rb}_2\text{Nb}_8\text{O}_{21}$	1315 (congruent)
$\text{Rb}_8\text{Nb}_{22}\text{O}_{59}$	1319 (congruent)
$\text{Rb}_2\text{Nb}_4\text{O}_{11}$	1185 (incongruent)
$\text{Rb}_4\text{Nb}_6\text{O}_{17}$	1148 (incongruent)
RbNbO_3	964 (incongruent)
$\text{Rb}_8\text{Nb}_6\text{O}_{19}$	830 (congruent)
$\text{Rb}_8\text{Nb}_2\text{O}_9$	1300 (congruent)

Crystals of $\text{Rb}_2\text{Nb}_4\text{O}_{11}$ suitable for X-ray examination were prepared by firing a mixture of carbonate and oxide, containing slight excess of the carbonate, over extended periods. The lattice was found to be monoclinic with

$$a = 12.95, b = 7.48, c = 14.92 \text{ \AA}, \beta = 106.4^\circ.$$

The space group was Cm , $C2$, or $C2/m$. The density measured (4.86 g.cm.^{-3}) shows rather poor agreement

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